**Acid Base Neutralization problems**

What volume of each of the following acids will react completely with 50.00ml of 0.200M NaOH?

1. .100 M HCl

2. 0.150 M HNO3

3. 0.200 M H3PO4

4. A 10.00 ml sample of vinegar, an aqueous solution of acid acid, (HC2H3O2) is titrated with 0.562M NaOH, and 16.58 mL of the NaOH is required to reach the equivalence point. What is the molarity of the vinegar? Acetic acid has one acidic hydrogen.

5. What volume of 0.0200M Calcium Hydroxide is required to neutralize 35.00 ml of 0.0500M Nitric Acid?

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